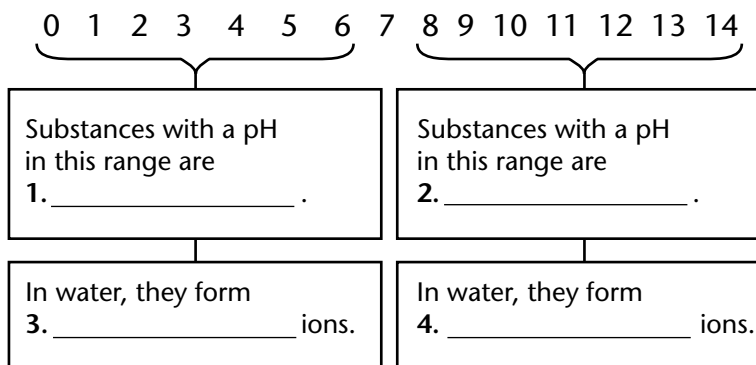


Acids, Bases, and Solutions ▪ *Review and Reinforce*

Acids and Bases in Solution

Understanding Main Ideas

Complete the concept map shown below and answer the following questions on a separate sheet of paper.



5. What is the difference between a strong acid and a weak acid?
6. What is the difference between a strong base and a weak base?
7. Which solution has a greater concentration of hydrogen ions (H^+), a solution with a pH of 3 or one with a pH of 7? Explain.
8. What are the products formed when a base reacts with an acid?
9. What is the pH of a neutral solution?

Building Vocabulary

Match each term with its definition by writing the letter of the correct symbol or definition on the line beside the term in the left column.

- | | |
|------------------------|--|
| ___ 10. hydroxide ion | a. ionic compound that can form from the reaction of an acid with a base |
| ___ 11. pH scale | b. reaction between an acid and a base |
| ___ 12. hydrogen ion | c. series of numbers that indicates the concentration of hydrogen ions in solution |
| ___ 13. neutralization | d. H^+ |
| ___ 14. salt | e. OH^- |