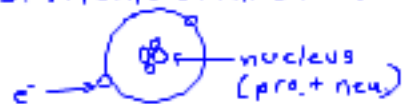


## After School Unit 2 Review

5. Atom: smallest unit



- a.  $H_2O$
- b.  $Na^+$
- c.  $C$
- d.  $O_2$

- e.  $NaCl$
- f.  $N$

molecule = 2 or more atoms chemically bonded



Two types of bonds

1. Ionic: Transfer of  $e^-$ s.  $Na^+Cl^-$
2. Covalent: Sharing of  $e^-$ s.  $H_2O$


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Ion is an atom w/ a (+) or (-) charge b/c of a gain or loss of electron(s)

Chem Rxn = Start w/ 1 set of chemicals that become a diff. set when bonds break + reform

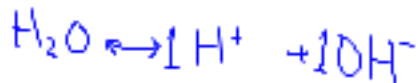


11. Cohesion: 2 molecules H-b's tied of same type 

Adhesion of diff. type   
↳ bandaid "glue" sticks to skin

(Big unit)  
Polymer is made of many monomers (single unit)

Train is made of many cars  
• pH (paper) measures the  $H^+$  in solution



Review (in class)

Format for test:

35 points

~ 20 m/c

~ 5 T/F

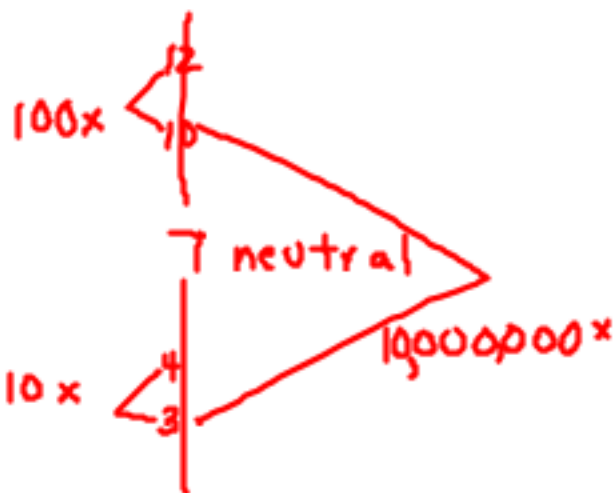
~ 8 short answer / f-i-t-b

~ 1 short essay

Polymer is made of monomers  
"Train" " " " " "Cars"

pH scale


14 basic



1 acidic

Ions:  $H^+$ ,  $Cl^-$ ,  $Na^+$

Atom: H, Cl, N

↳  = 1 atom  
of the element  
Helium.

5. C = atom  
 $NaCl^-$  = compound (ionic)  
transfer of  $e^-$



pH scale measures # of  
in a solution

$H_2O$  = molecule  
(no ions)  
↳ shares electrons  
covalent bond!  
another ex:  $O_2$

Isotopes

12  
p=6  
n=6



14  
p=6  
n=8



13 Mixture (2 solids)  
not chemically bonded

Suspension = (1 liquid +  
1 non-dissolvable solid)

20 Neutral:  $\text{OH}^- = \text{H}^+$   
Acids have more  $\text{H}^+$  than  $\text{OH}^-$   
Bases have more  $\text{OH}^-$  than  $\text{H}^+$