

# Evaporation and Intermolecular Attractions

In this experiment, Temperature Probes are placed in various liquids. Evaporation occurs when the probe is removed from the liquid's container. This evaporation is an endothermic process that results in a temperature decrease. The magnitude of a temperature decrease is, like viscosity and boiling temperature, related to the strength of intermolecular forces of attraction. In this experiment, you will study temperature changes caused by the evaporation of several liquids and relate the temperature changes to the strength of intermolecular forces of attraction. You will use the results to predict, and then measure, the temperature change for several other liquids.

You will encounter two types of organic compounds in this experiment—alkanes and alcohols. The two alkanes are n-pentane,  $C_5H_{12}$ , and n-hexane,  $C_6H_{14}$ . In addition to carbon and hydrogen atoms, alcohols also contain the -OH functional group. Methanol,  $CH_3OH$ , and ethanol,  $C_2H_5OH$ , are two of the alcohols that we will use in this experiment. You will examine the molecular structure of alkanes and alcohols for the presence and relative strength of two intermolecular forces—hydrogen bonding and dispersion forces.

## OBJECTIVES

In this experiment, you will

- Study temperature changes caused by the evaporation of several liquids.
- Relate the temperature changes to the strength of intermolecular forces of attraction.

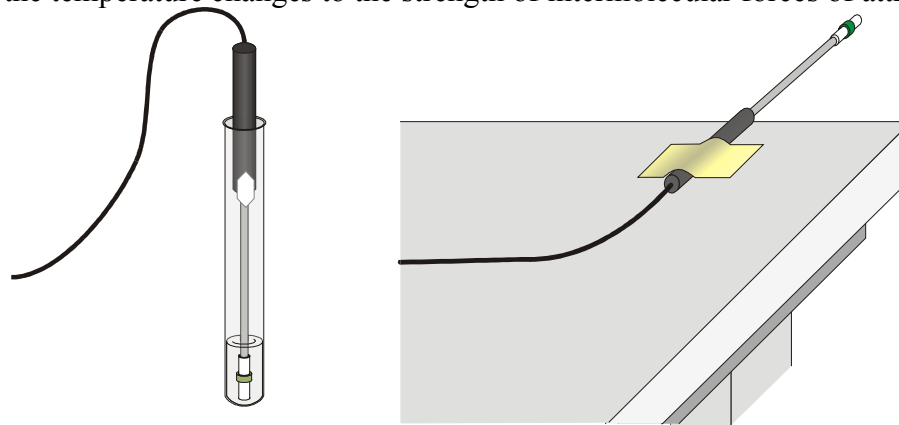


Figure 1

## MATERIALS

computer  
Vernier computer interface  
LoggerPro  
two Temperature Probes  
6 pieces of filter paper (2.5 cm × 2.5 cm)  
2 small rubber bands  
masking tape

methanol (methyl alcohol)  
ethanol (ethyl alcohol)  
1-propanol  
1-butanol  
n-pentane  
n-hexane

## PRE-LAB EXERCISE

Prior to doing the experiment, complete the Pre-Lab table. The name and formula are given for each compound. Draw a structural formula for a molecule of each compound. Then determine the molecular weight of each of the molecules. Dispersion forces exist between any two molecules, and generally increase as the molecular weight of the molecule increases. Next, examine each molecule for the presence of hydrogen bonding. Before hydrogen bonding can occur, a hydrogen atom must be bonded directly to an N, O, or F atom within the molecule. Tell whether or not each molecule has hydrogen-bonding capability.

## PROCEDURE

1. Obtain and wear goggles! **CAUTION:** The compounds used in this experiment are flammable and poisonous. Avoid inhaling their vapors. Avoid contacting them with your skin or clothing. Be sure there are no open flames in the lab during this experiment. Notify your instructor immediately if an accident occurs.
2. Connect the probes to the computer interface. Prepare the computer for data collection by opening the file "09 Evaporation" from the *Chemistry with Vernier* folder.
3. Wrap Probe 1 and Probe 2 with square pieces of filter paper secured by small rubber bands as shown in Figure 1. Roll the filter paper around the probe tip in the shape of a cylinder. Hint: First slip the rubber band up on the probe, wrap the paper around the probe, and then finally slip the rubber band over the wrapped paper. The paper should be even with the probe end.
4. Stand Probe 1 in the ethanol container and Probe 2 in the 1-propanol container. Make sure the containers do not tip over.
5. Prepare 2 pieces of masking tape, each about 10 cm long, to be used to tape the probes in position during Step 6.
6. After the probes have been in the liquids for at least 30 seconds, begin data collection by clicking . Monitor the temperature for 15 seconds to establish the initial temperature of each liquid. Then simultaneously remove the probes from the liquids and tape them so the probe tips extend 5 cm over the edge of the table top as shown in Figure 1.
7. When both temperatures have reached minimums and have begun to increase, click  to end data collection. Click the Statistics button,  , then click  to display a box for both probes. Record the maximum ( $t_1$ ) and minimum ( $t_2$ ) values for Temperature 1 (ethanol) and Temperature 2 (1-propanol).
8. For each liquid, subtract the minimum temperature from the maximum temperature to determine  $\Delta t$ , the temperature change during evaporation.
9. Roll the rubber band up the probe shaft and dispose of the filter paper as directed by your instructor.
10. Based on the  $\Delta t$  values you obtained for these two substances, plus information in the Pre-Lab exercise, *predict* the size of the  $\Delta t$  value for 1-butanol. Compare its hydrogen-bonding capability and molecular weight to those of ethanol and 1-propanol. Record your predicted  $\Delta t$ , then explain how you arrived at this answer in the space provided. Do the same for n-pentane. It is not important that you predict the exact  $\Delta t$  value; simply estimate a logical value that is higher, lower, or between the previous  $\Delta t$  values.

11. Test your prediction in Step 10 by repeating Steps 3-9 using 1-butanol for Probe 1 and n-pentane for Probe 2.
12. Based on the  $\Delta t$  values you have obtained for all four substances, plus information in the Pre-Lab exercise, predict the  $\Delta t$  values for methanol and n-hexane. Compare the hydrogen-bonding capability and molecular weight of methanol and n-hexane to those of the previous four liquids. Record your predicted  $\Delta t$ , then explain how you arrived at this answer in the space provided.
13. Test your prediction in Step 12 by repeating Steps 3-9, using methanol with Probe 1 and n-hexane with Probe 2.

## **PROCESSING THE DATA**

1. Two of the liquids, n-pentane and 1-butanol, had nearly the same molecular weights, but significantly different  $\Delta t$  values. Explain the difference in  $\Delta t$  values of these substances, based on their intermolecular forces.
2. Which of the alcohols studied has the strongest intermolecular forces of attraction? The weakest intermolecular forces? Explain using the results of this experiment.
3. Which of the alkanes studied has the stronger intermolecular forces of attraction? The weaker intermolecular forces? Explain using the results of this experiment.
4. Plot a graph of  $\Delta t$  values of the four alcohols versus their respective molecular weights. Plot molecular weight on the horizontal axis and  $\Delta t$  on the vertical axis.

## PRE-LAB

Substance	Formula	Structural Formulas	Molecular Weight	Hydrogen Bond (Yes or No)
ethanol	C <sub>2</sub> H <sub>5</sub> OH			
1-propanol	C <sub>3</sub> H <sub>7</sub> OH			
1-butanol	C <sub>4</sub> H <sub>9</sub> OH			
n-pentane	C <sub>5</sub> H <sub>12</sub>			
methanol	CH <sub>3</sub> OH			
n-hexane	C <sub>6</sub> H <sub>14</sub>			

## DATA TABLE

Substance	t <sub>1</sub> (°C)	t <sub>2</sub> (°C)	Δt (t <sub>1</sub> -t <sub>2</sub> ) (°C)	Predicted Δt (°C)	Explanation
ethanol					
1-propanol					
1-butanol					
n-pentane					
methanol					
n-hexane					